

	PREDICTION	OBSERVATION
Atomic weight . . . . .	72.0	72.3
Specific gravity . . . . .	5.5	5.469
Atomic volume . . . . .	13	13.2
Formula of oxide . . . . .	GeO <sub>2</sub>	GeO <sub>2</sub>
Specific gravity of oxide . . . . .	4.7	4.703
Formula of chloride . . . . .	GeCl <sub>4</sub>	GeCl <sub>4</sub>
Boiling point of chloride . . . . .	Less than 100°	86°
Specific gravity of chloride . . . . .	1.9	1.9
Formula of fluoride . . . . .	GeFl <sub>4</sub>	GeFl <sub>4</sub>
Formula of ethyl compound . . . . .	Ge(C <sub>2</sub> H <sub>5</sub> ) <sub>4</sub>	Ge(C <sub>2</sub> H <sub>5</sub> ) <sub>4</sub>
Specific gravity of ethyl compound	0.96	Lower than water

Finally it is to be especially noted that, upon arranging the known elements in a table rationally constructed upon the basis of the above recorded facts, comparatively few spaces within the range of known atomic weights remain to be filled. The conclusion is obvious that very few elements now unknown are possible unless they possess very high atomic weights. But the apparent transmutation of radium into helium is a pretty clear indication that elements of very high atomic weight may be unstable. If they have existed in number and large quantity, they probably have long since ceased so to exist, except perhaps in the interior of celestial bodies, and they are not likely elsewhere to complicate natural phenomena by their unknown properties.