the high latent heats of fusion and evaporation of water upon the meteorological processes. When, for example, a body of water becomes cooled to its freezing point, the further abstraction of heat cannot lower its temperature below that point, which, to be sure, is somewhat variable in the case of salt water. And so long as water and ice exist in contact, the system constitutes a thermostat, a very accurate one if the water be fresh, which changes only in respect to the quantities of ice and water as heat is added or removed.[1] Heating serves merely to melt the ice, cooling to freeze the water. Accordingly, as long as the earth shall remain habitable the cooling of its oceans and seas will remain rigidly limited by their freezing point. However inclement the atmosphere, the ocean can always support life until the final extinction of water by cold. It is worthy of note that the freezing point of water, though to man with his carefully regulated body temperature apparently low, is in reality very high indeed compared with that of any like substances, — perhaps 100° centigrade above the average.

[1] In fact, there is no better means of obtaining a constant temperature in the chemical laboratory than by mixing pure ice with pure water.