

TABLE OF MELTING POINTS

		DEGREES
Water . . . . .	H <sub>2</sub> O	0
Hydride of antimony . . . . .	SbH <sub>3</sub>	-91.5
Hydride of arsenic . . . . .	AsH <sub>3</sub>	-113.5
Hydrobromic acid . . . . .	HBr	-87
Hydrochloric acid . . . . .	HCl	-112.5
Hydrofluoric acid . . . . .	HF	-92.3
Hydriodic acid . . . . .	HI	-50
Methane . . . . .	CH <sub>4</sub>	-185.8
Carbon dioxide . . . . .	CO <sub>2</sub>	-57
Hydride of phosphorus . . . . .	PH <sub>3</sub>	-132.5
Hydrogen sulphide . . . . .	H <sub>2</sub> S	-85.6
Sulphurous oxide . . . . .	SO <sub>2</sub>	-72.7
Ammonia . . . . .	NH <sub>3</sub>	-75
Nitric oxide . . . . .	NO	-167

This is, no doubt, one of the most important facts with which we are concerned, for while a very large number of chemical processes take place quite freely at 0°, the conditions are very different at the freezing point of ammonia, for instance. At that temperature the velocity of most chemical processes is but a fraction of one per cent of their velocity at 0°, and a large part of the chemical activity which is familiar to us ceases.

The result of the unusually high freezing point of water and of the phenomenon of latent heat is felt, however, not merely in the avoidance of an excessive fall in the temperature of lakes and seas. As above explained,