Substance	Formula	Melting Point	Latent Heat of Fusion
Water	H_2O	0°	80 Calories
Ammonia	NH_3	−75	108
Antimony chloride .	$SbCl_3$	73	13.4
Antimony bromide .	$SbBr_3$	94	9.7
Lead chloride . .	$PbCl_2$	485	20.9
Calcium chloride .	$CaCl_2 \cdot 6\,H_2O$	28.5	40.7
Potassium nitrate .	KNO_3	339	47.4
Sodium nitrate . .	$NaNO_3$	310.5	63
Phosphoric acid .	H_3PO_4	18	25.7
Nitric acid . . .	HNO_3	−47°	9.5
Sulphuric acid . .	H_2SO_4	10.3	24.
Sulphuric oxide . .	SO_3		76.7
Ethylene bromide .	$C_2H_4Br_2$	8	13
Formic acid . . .	$H.COOH$	−7.5	57.4
Chloral hydrate .	$C_2H_3Cl_3O_2$	46	33.2
Dimethyl oxalate .	$C_2O_4(CH_3)_2$	49.5	42.6
Acetic acid . . .	CH_3COOH		43.7
Glycerine	$C_3H_8O_3$	13	42.5
Stearic acid . . .	$C_{18}H_{36}O_2$	64	47.6
Benzene	C_6H_6	5.3	30.1
Nitrobenzene . .	$C_6H_5NO_2$	−9.21	22.3
Di-chlorbenzene .	$C_6H_4Cl_2$	52.5	29.9
p-Toluidine . . .	C_7H_9N		35.8
Phenol	C_6H_5OH	25.4	24.9
Menthol	$C_{10}H_{20}O$	42	18.9
Phenylhydrazine .	$C_6H_5.NH.NH_2$		24.5
Phenylacetic acid .	$C_6H_5.CH_2.COOH$	75	25.4
Naphthaline . . .	$C_{10}H_8$	80	35.7

Accordingly, the processes above described possess nearly the highest possible efficiency. A very large amount of heat must be abstracted from a body of water before it can be solidified; after a given amount of cool-