

This is the case because the tendency of the electrically charged ions to reunite and form electrically neutral molecules must be less the greater the dielectric constant of the solvent. Now the dielectric constant of water is nearly the highest at present known, and therefore ionization in water is on that account also more extensive than in almost any other solvent.

Finally, for reasons that are not yet understood, the process of ionization in other solvents than water is a much more complex affair, and there can be no doubt that such complexity limits the phenomena which are dependent upon ionization.<sup>1</sup>

electrostatic attraction between the ions is alone weakened, and hence the degree of dissociation is increased." — LE BLANC, "A Textbook of Electro-Chemistry." New York, 1907, p. 147.

<sup>1</sup> "It would be natural to expect that the conceptions which have been found serviceable in the case of solutions in water could be applied directly to solutions in other solvents, keeping in mind that, according to the individual nature of any given solvent, the degree of dissociation, the migration velocity of the ions, and consequently the conductivity of a solution of a given concentration would be different. It is a noteworthy fact, however, that the behavior of non-aqueous is much more complicated than that of aqueous solutions. This is shown especially by the investigation of the conductivity of solutions of various substances in liquid sulfur dioxide made by Walden and Centnerszwer. Neither the law of the independent migration of the ions, nor the law that by increas-