http://www.geology.19thcenturyscience.org/books/1913-Henderson-Fitness/README.htm

tion of 1 kilogram of carbon dioxide in 100 liters of water, to which sodium hydrate is being added. At the beginning of the experiment the hydrogen ion concentration will be approximately 0.0001 N, almost exactly one thousand times that at neutrality, and the hydroxyl ion concentration 0.0000000001 N, one one-thousandth that at neutrality. If, now, the sodium hydrate be added to the solution in successive portions, the change of reaction will be as indicated in the following table: —

Amount of NaOH	$\overset{+}{(H)}$	$\overset{-}{(OH)}$	Intensity compared with a Neutral Solution of	
			Acidity	Alkalinity
Grams				
0	0.0001 N	0.0000000001 N	1000	0.001
50	0.0000052	0.000000002	52	0.02
100	0.0000024	0.000000004	24	0.04
150	0.0000015	0.000000006	15	0.06
200	0.0000011	0.000000009	11	0.09
250	0.0000008	0.000000012	8	0.12
300	0.0000006	0.000000017	6	0.17
350	0.0000005	0.000000020	5	0.20
400	0.0000004	0.000000025	4	0.25
450	0.0000003	0.000000033	3	0.33
500	0.00000025	0.00000004	2.5	0.4
550	0.00000020	0.00000005	2.0	0.5
600	0.00000015	0.00000007	1.5	0.7
650	0.00000012	0.00000008	1.2	0.8
700	0.00000009	0.00000011	0.9	1.1
750	0.00000006	0.00000017	0.6	1.7
800	0.00000004	0.00000025	0.4	2.5
850	0.00000002	0.0000005	0.2	5.0